淡江大學八十九學年度日間部轉學生招生考試試題

系別:化學系三年級

科目:分析化學

本試題共 1 頁

- —.Define the following terms:
 - (1) Buffer capacity
 - (2) Parts per million (ppm)
 - (3) Volhard method

(30%)

- (4) Isoelectric point
- (5) Kjeldahl method
- (6) Chelating agent
- =.(a)What mass of AgNO₃(169.9 g/mol)is needed to convert 2.33 g of Na₂CO₃(106.0 g/mol) to Ag₂CO₃? (b)What mass of Ag₂CO₃(275.7 g/mol) will be formed? (10%)
- \equiv . Calculate the standard deviation of the result of $\frac{4.10(\pm 0.02) \times 0.0050(\pm 0.0001)}{1.97(\pm 0.04)} = 0.0104(\pm ?)$
- 四.Calculate the solubility of $Ba(IO_3)_2$ in a solution prepared by mixing 200 mL of 0.0100 M $Ba(NO_3)_2$ with 100 mL of 0.100 M $NaIO_3$. (Note: Ksp for $Ba(IO_3)_2$ is 1.57×10^{-9})
- 五.Calculate the hydronium ion concentration in 0.120 M nitrous acid. The principal equilibrium is

 $HNO_2 + H_2O \rightleftharpoons H_3O^+ + NO_2^-$ (Note: Ka for nitrous acid is 7.1×10^{-4})

六.Calculate the thermodynamic potential of the following cell and indicate whether it is galvanic or electrolytic. (10%) Pt | UO₂²⁺(0.0150 M), U⁴⁺(0.200 M), H⁺(0.0300 M) || Fe²⁺(0.0100

M), Fe^{3+} (0.0250 M) | Pt

The two half-reactions are

$$Fe^{3+} + e^{-} = Fe^{2+}$$
 $E^{0} = +0.771 \text{ V}$ $UO_{2}^{2+} + 4H^{+} + 2e^{-} = U^{4+} + 2H_{2}O$ $E^{0} = +0.334 \text{ V}$

七. (a)何謂 C₁₈ 逆相(reversed-phase HPLC)分離管柱? (b)C₁₈ 分離管柱主要用於分離那類化合物? (10%)

八.說明毛細管電泳(Capillary Electrophoresis)分離法的原理.(5%)

九. (a)何謂超臨界流體層析法(SFC)?

(b)和氣體層析法(GC)比較,那些溶質比較適用於SFC來分析?

(10%)